

Chapter 10 Chemical Quantities Assessment Answers

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Chapter 10 Chemical Quantities Assessment

Chapter 10 Chemical Quantities Assessment Section 10.1 - The Mole: A Measurement of Matter You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance. 6.02×10^{23} is called Avogadro's number.

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Chapter 10 - Chemical Quantities - 10 Assessment - Page 342: 121 Answer An atom is the smallest particle of an element that retains its identity in a chemical reaction, whereas a molecule is

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a group of atoms joined together by covalent bonds.

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Section 10.3 - Percent Composition and Chemical Formulas. The percent by mass (percent composition) of an element in a compound is the number of grams of the element divided by the mass in grams of the compound multiplied by 100%. % mass of element = $\frac{\text{mass of element}}{\text{mass of compound}} \times 100$.

Chapter 10 - Chemical Quantities

Chapter 10 Chemical Quantities 241 Section Review Objectives • Relate Avogadro's number to a mole of a substance • Calculate the mass of a mole of any substance • Describe methods of measuring the amount of something • Compare and contrast the atomic mass of an element and its molar mass Vocabulary Key Equations • moles representative particles • representative particles moles.

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Chapter 10 Chemical Quantities⁹¹ SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296)

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10 1 kg 103 g 368 Chapter 10 Chemical Calculations and Chemical Equations. 10.1 Equation Stoichiometry 369 The ratio of moles of P 4O 10 to moles of P (which came from the subscripts in the chemical formula, P 4O 10) provided the key conversion factor that allowed us to convert

Chapter 10 Chemical Calculations and equations

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