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A solution is made so that it is 2.5 M  $\text{Ca}(\text{NO}_3)_2$ . Therefore the concentration of  $\text{Ca}^{2+}$  is 2.5 M and the concentration of  $\text{NO}_3^-$  is 5.0 M. The total concentration of all particles is 7.5

M.Explain.Since each mole of  $\text{Ca}(\text{NO}_3)_2$  breaks up into 1 mole of  $\text{Ca}^{2+}$ , the concentration of  $\text{Ca}^{2+}$  would be the same as the concentration of  $\text{Ca}(\text{NO}_3)_2$ .

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Information: Molarity Concentration is a term that describes the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions contain only a little.

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concentration in units of molarity by following these steps: a) Convert the grams of NaCl to moles of NaCl. b) Calculate the litres of solution by dividing the given millilitres by 1000. c) Divide moles by the litres of solution. 4. Calculate the concentration (in units of molarity) of the following solutions: a) 42.1 g of Ca(NO<sub>3</sub>)<sub>2</sub>

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